

Full Marks: 63.5 + 1.5

Pass Marks: 32

Time Duration: 1 hour 15 mins

classmate

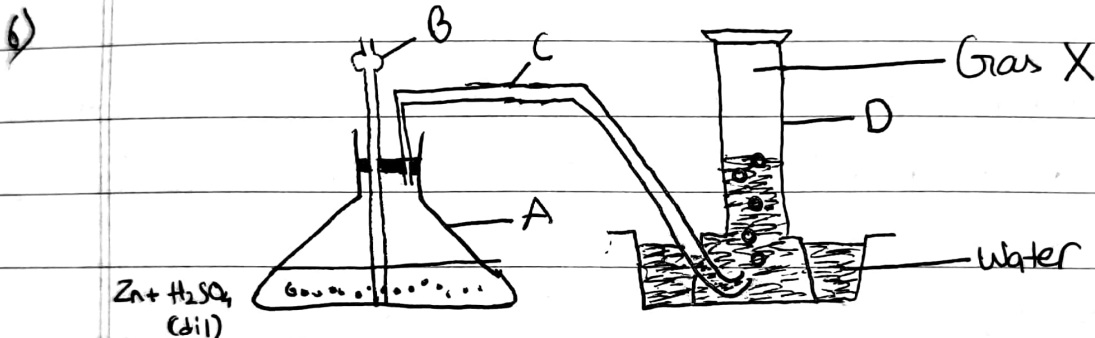
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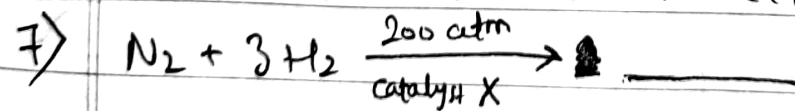
Chemistry Test Ch-7, 8

7/2/21

- 1) Hydrogen was discovered by _____ (1)
- 2) Hydrogen is the lightest element on the earth TIF (1)
- 3) Sodium reacts with hydrogen to form _____ gas. (1)
- 4) Why the mixture of oxygen and hydrogen is used for welding. (1)
- 5) What happens when steam is passed over carbon. Give the chemical equation for the reaction as well. (2)



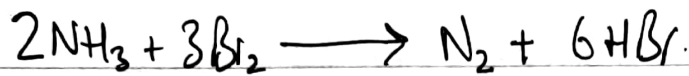
- a) ~~What is the aim~~ Mark A-D (2)
- b) What is the aim of the test. (1)
- c) ~~What~~ flow is ~~the~~ gas X collected by looking at the setup? (1)
- d) Why is a dilute ^{acid} used? (1)
- e) Write a chemical eq for the experiment. (2)



- a) Name the process. (1)
- b) What is the product formed? (1)
- c) What is catalyst X? (1)

8) What is hydrogenation? (2)

9)



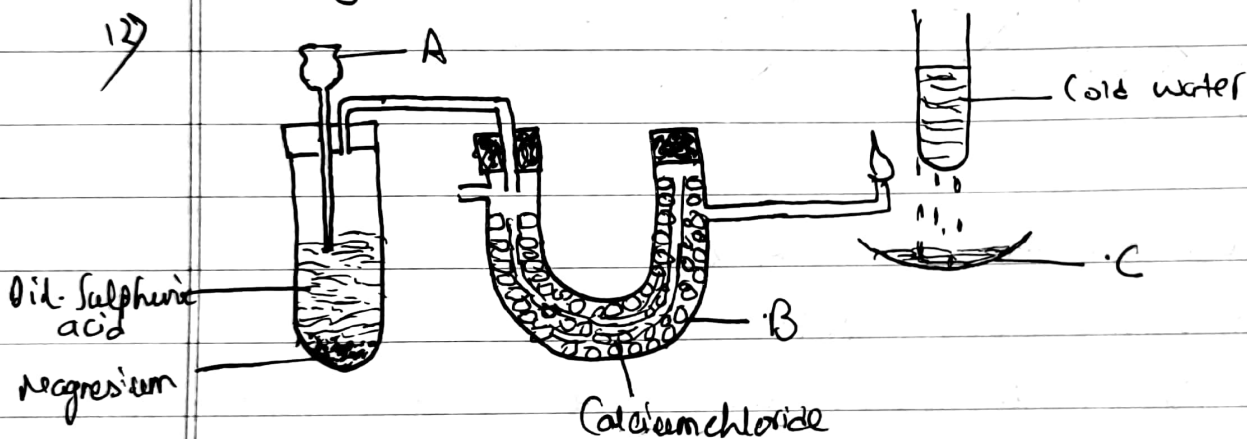
- What type of reaction is this? (1)
- Ammonia is _____ agent in the above reaction. (1)
- Which other element is the same agent as ammonia? (1)
- If bromine is an oxidising agent that means ~~that~~ bromine is (1)
reduced. T/F.

10)

How is methane formed? Write a chemical eq for it. (3)

- "Hydrogen reduces metal oxides". Is this statement true? Why? (3)

12)



- Mark A-C (1 1/2)
- What is the aim of the experiment? (1)
- Write a chemical eq for the above reaction. (2)

11) Tritium has 2 neutrons. T/F (1)

12) A metal X reacts with steam to form an oxide of the metal on its surface which prevents further reaction.



Name the metal.

- 13) When calcium is added to water a solution is formed along with the release of a gas. A pop sound is heard by bringing a burning splint to the mouth of the test tube. When we dip a red litmus paper to the solution it turns blue. What conclusion can you draw out from this experiment. Write a chemical equation for the same. (3)
- 14) What is the difference between temporary hardness and permanent hardness? How can we treat them? (4)
- 15) What is water of crystallisation? How can we make anhydrous crystals of copper sulphate? (3)
- 16) What are deliquescent substances? Give names of three deliquescent substances. (3)
- 17) Hygroscopic substances are used for absorbing water and forming solutions. T/F. (1)
- 18) Show the impact of heat on copper sulphate crystals with the help of a chemical reaction. (2)
- 19) "Magnesium reacts with water more readily than calcium". Is this statement true? Why? (2)
- 20) What is super saturated solution? What are the factors on which solubility depends? (4)
- 21) When a non-metallic oxide reacts with water _____ is formed. Give a chemical reaction for the same. (3)
- 22) Electrolytes are substances that disperse at molecular level ~~in a solution~~ in water because they are covalent compounds. T/F (1)
- 23) Complete and balance the eq $\text{Ca}(\text{HCO}_3)_2 + \text{Na}_2\text{CO}_3 \rightarrow \bullet \text{_____} \text{_____}$ (2)